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REVIEW Stoichiometry SECTION 3 PROBLEMS Write
the answer on the line to the left. Show all your work
in the space provided. 1. _____ The actual yield of a
reaction is 22 g and the theoretical yield is 25 g.
Calculate the percentage yield. 2. 6.0 mol of N₂ are
mixed with 12.0 mol of H CHAPTER 9 REVIEW
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the mass relationships of elements in c... Calculations involving the mass relationships between reactant... A conversion factor that relates the amounts in moles of any t... The substance that controls the quantity of product that can f ...

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CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction. (d) number of valence electrons involved in a reaction ...

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chemical reaction. 3. Definitions. A mole ratio is ...

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mass relationships between reactants and products in
a chemical reaction. Reaction stoichiometry, the
subject of this chapter, is based on chemical

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